

Vocabulary for Pure Substances and Mixtures Unit ANSW

Date: _____

Agitate - To stir, shake or otherwise mix

Atom - the smallest unit of matter composed of protons, neutrons and electrons.

Change – Chemical - a change that is not easily reversed. It changes the configurations of the molecules involved in the change.

Change – Physical - A phase shift due to the addition or removal of thermal energy (heat).

Concentration - The relative amount of solute to solvent in a solution. Often stated as a percentage, g/L or ppm.

Dilute - A solution in which there is very little solute to solvent (low concentration).

Evaporate – The process by which a liquid phase becomes a gas phase. It requires the addition of heat.

Gas – The state of matter which has the highest amount of energy. It is free to fill the container it is in.

Liquid - The state of matter that is between solid and gas. It is capable of flowing and fills the bottom of its container.

Mechanical Mixture – A substance that has two or more definite parts. It does not become a solution.

Molecule - A pure substance that is composed of a defined combination of atoms. H_2O , CO_2 and CH_4 are small molecules.

Properties – Characteristic aspects or behaviours of substances. A characteristic property of metals is conduction of electricity.

Pure – without impurities or other parts.

Saturated – A solution that has the maximum amount of solute dissolved at that temperature in a given solvent. Sometimes known as the saturation point.

Solid – The phase of matter in which the particles are bound rigidly together. This phase keeps its shape and is low energy.

Solute – The part of a solution that is being dissolved. It is the smaller component in the solution.

Solution – A mixture of two or more different substances that is the same throughout

Solvent – The part of a solution doing the dissolving – it is the larger component in the solution.

Supersaturated – Having more dissolved solute than should normally be possible. This is rare since excess solute usually crystallizes and drops out of solution.

Unsaturated – The state in which a solution can dissolve more solute into solution. (It is below the saturation point)